

Synthesis and Electrochemical Analysis for NiMn_2O_4 as Anode Material in Lithium-Ion Batteries

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by

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I, Sweta Mehra (23/MSCCHE/59), hereby declare that the work presented in the dissertation entitled **“Synthesis and Electrochemical Analysis for NiMn₂O₄ as Anode Material in Lithium-Ion Batteries”** is an original work that has been carried out by me during the period from 2024 to 2025 at Lithium-Ion Battery Technology Laboratory (LIBT), Delhi Technological University under the guidance of Dr. Anil Kumar and Dr. Amrish K. Panwar.

The research presented in this dissertation is an authentic record of my work and has been submitted in partial fulfilment of the requirements for the award of the Degree of Master of Science in Chemistry to the Department of Applied Chemistry, Delhi Technological University, New Delhi, India.

I have not submitted the matter presented in this project for the award of any other degree or diploma of any other Institute or University.

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This is to certify that Sweta Mehra (23/MSCCHE/59) have carried out her research work presented in this thesis entitled **“Synthesis and Electrochemical Analysis for NiMn₂O₄ as Anode Material in Lithium-Ion Batteries”** for the award of Degree of Master of Science in Chemistry from Department of Applied Chemistry, Delhi Technological University, Delhi, under our supervision. The dissertation embodies the results of original work, and studies are carried out by herself, and the content of the thesis does not form the basis of the award of any other degree to the candidate or to any other body else from this or any other University/Institution.

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ABSTRACT

Lithium-Ion Batteries (LIBs) have become a rising energy storage solution for various applications, ranging from compact and portable electronics to electric vehicles. Since, the demand for these rechargeable batteries is increasing day by day. Hence, there is a need for new and highly efficient electrode and electrolyte materials to enhance the overall performance of LIBs. Transition metals, known for their ability to exhibit variable oxidation states, are ideal candidates for anode materials. Among these, mixed transition metal oxides (MTMOs) have gained significant attention for their potential to improve battery performance. In this study, spinel type NiMn_2O_4 (NMO) has been synthesised using a citric acid-assisted solid-state method. The X-ray diffraction patterns of the synthesized materials confirm the formation of the proper phase with $\text{Fd}3\text{m}$ cubic crystal symmetry. The SEM and TEM analysis revealed the particle size and morphology of NMO, showing well-defined hexagonal-shape structure. The electrochemical properties such as Li-ion charge transfer resistance, redox potentials, charge discharge capacities cyclability and Li-ion diffusion coefficient of NMO in Li half-cell has been measured using Electrochemical Impedance Spectroscopy (EIS), Cyclic Voltammetry (CV), and Galvanostatic Charge-Discharge (GCD).



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CONTENTS

CANDIDATE DECLARATION	I
CERTIFICATE.....	II
PLAGIARISM VERIFICATION.....	III
ABSTRACT.....	IV
ACKNOWLEDGEMENT.....	V
CONTENTS	VI
LIST OF FIGURES.....	VIII
LIST OF TABLES.....	IX
ABBREVIATIONS	X

Chapter 1: INTRODUCTION..... 1

1.1 Rechargeable Batteries1

1.2 Lithium-Ion Batteries2

1.2.1 Components of LIBs.....2

1.2.1.1 Electrolyte..... 3

1.2.1.2 Cathode.....4

1.2.2.3 Separator..... 5

1.2.2.4 Anode6

1.3 Principle and Working7

1.3.1 Electrochemistry.....7

1.3.2 Electrochemical Performance.....8

1.3.3.1 Charge/Dis-charge9

1.3.3.2 Specific Capacitance.....9

1.3.3.3 Lithium-Ion Diffusion Coefficient9

Chapter 2: LITERATURE REVIEW 10

2.1 Types of Anodes.....10

2.1.1 Transition Metal Oxides12

2.1.2 Mixed Transition Metal Oxides13

2.1.2.1 Nickel Manganese Oxide..... 13

Chapter 3: EXPERIMENTAL PROCEDURE	16
3.1 Materials and Methods	16
3.2 Material Characterization	17
3.2.1 X-ray diffraction (XRD).....	17
3.2.1 Scanning Electron Microscopy (SEM).....	18
3.2.1 Energy Dispersive x-ray Spectroscopy (EDS)	18
3.2.1 Transmission Electron Microscopy (TEM).....	19
3.2.1 Selected Area Electron diffraction (SAED)	19
3.3 Electrode Fabrication.....	20
3.4 Electrochemical Measurement	21
3.4.1 Electrochemical Impedance Spectroscopy (EIS)	21
3.4.2 Cyclic Voltammetry (CV)	22
3.4.3 Galvanostatic Charge/Dis-charge (GCD).....	22
Chapter 4: RESULTS AND DISCUSSIONS.....	23
4.1 Phase Analysis.....	23
4.2 Morphology Analysis	24
4.3 Electrochemical Analysis	28
4.3.1 EIS	28
4.3.2 CV	29
4.3.3 GCD.....	30
Chapter 5: CONCLUSIONS.....	32
5.1 Conclusion	32
5.2 Future Scopes.....	32
REFERENCES	33
APPENDICES	i
List of Publications.....	i
List of Conferences Attended.....	ii
Plagiarism Report.....	iv

LIST OF FIGURES

- Fig. 1** Components of Lithium-Ion Battery
- Fig. 2** Dis-charging process of LIBs
- Fig. 3** Charging process of LIBs
- Fig. 4** Structure analysis of NMO, a spinel AB_2O_4 -type material
- Fig. 5** Schematic representation of synthesis process
- Fig. 6** Diagrammatic representation of XRD
- Fig. 7** Diagrammatic representation of SEM
- Fig. 8** Diagrammatic representation of TEM
- Fig. 9** Coin-cell assembly representation
- Fig. 10** Diagrammatic Representation of Potentiostat for Electrochemical analysis
- Fig. 11** X-ray analysis of NMO calcined at 800°C for 3 hours
- Fig. 12** Williamson-Hall (W-H) plot of synthesized sample.
- Fig. 13** SEM analysis of synthesised NMO at a) 50k x, b) 100k x, c) 120k x, and elemental mapping images of d) Oxygen (green), e) Nickel (red), f) Manganese (yellow) and g) EDS analysis along with map sum spectrum
- Fig. 14** SEM histogram for average particle size
- Fig. 15** TEM analysis of NMO at a) 200nm, b) 50 nm, and c) SAED analysis at 2 nm range
- Fig. 16** TEM histogram for average particle size
- Fig. 17** Nyquist plot showing impedance conducted at 1 MHz to 100 Hz frequency range and 5 mV AC amplitude
- Fig. 18** Cyclic voltammetry curves obtained at 0.1 mV s⁻¹ scan rate
- Fig. 19** GCD plot showing charge and discharge capacity at 0.01C-rate

LIST OF TABLES

Table 1	Analysis of Cathode Materials
Table 2	Analysis of Anode Materials
Table 3	Common TMO anodes along with their advantages and dis-advantages

ABBREVIATIONS

LIBs	Lithium-Ion Batteries
SEI	Solid-Electrolyte Interface
NMO	Nickel Manganese Oxide (NiMn_2O_4)
LTO	Lithium Titanate Oxide
TMOs	Transition Metal Oxides
MTMOs	Mixed Transition Metal Oxides
XRD	X-ray Diffraction
SEM	Scanning Electron Microscopy
EDS	Energy Dispersive X-ray Spectroscopy
TEM	Transmission Electron Microscopy
SAED	Selected Area Electron diffraction
EIS	Electrochemical Impedance Spectroscopy
CV	Cyclic Voltammetry
GCD	Galvanostatic Charge/Dis-charge
EC	Ethylene carbonate
DMC	Dimethyl carbonate

Chapter 1

INTRODUCTION

1.1 Rechargeable Batteries

Rechargeable batteries, also known as secondary batteries, are energy storage devices that can be repeatedly charged and discharged, unlike primary (non-rechargeable) batteries [1]. Their ability to efficiently store and deliver electrical energy makes them essential for a variety of applications, from electrical to energy storage technologies. The evolution of rechargeable battery technology has been driven by the increasing demand for higher energy density, longer cycle life, improved safety, and lower cost [2-3].

Traditionally, several types of rechargeable batteries have played significant roles:

- **Lead-Acid Batteries:** The Lead-Acid based rechargeable battery are known for their robustness, low cost, and ability to deliver high surge currents, making them suitable for automotive starting, lighting, and ignition (SLI) applications, as well as uninterruptible power supplies (UPS) [4]. However, their relatively low energy density and environmental concerns regarding lead disposal limit their application in portable devices [5].
- **Nickel-Cadmium (Ni-Cd) Batteries:** Developed in the early 20th century, Ni-Cd batteries offered better energy density and cycle life compared to lead-acid [6]. They are widely used in many applications. A major drawback is the "memory effect," where repeated partial discharge cycles can lead to a reduction in capacity. Furthermore, cadmium is a highly toxic heavy metal, leading to their phasing out in many regions [7-8].
- **Nickel-Metal Hydride Batteries:** These batteries replaced Ni-Cd batteries as a more green alternative solution. Ni-MH batteries have a higher energy density than Ni-Cd and exhibit a

less pronounced memory effect [9]. While an improvement, their energy density and voltage are still lower than those of lithium-ion batteries [10].

- **Lithium-Ion Batteries (LIBs):** Emerged in the early 1990s and revolutionized portable electronics due to their high energy density, high operating voltage, and absence of memory effect [11]. They quickly became the dominant technology for smartphones, laptops, and more recently, electric vehicles (EVs) and energy storage devices. Their continuous development aims to enhance safety, further increase energy density, extend cycle life, and reduce cost [12-13].
- **Other Emerging Technologies:** Research and development are ongoing for various next-generation battery technologies, including solid-state batteries, lithium-sulfur (Li-S) batteries, lithium-air (Li-air) batteries, and sodium-ion (Na-ion) batteries, which aim to surpass the performance limitations of current Li-ion technology in terms of energy density, safety, and cost [14-15].

Each type of rechargeable battery possesses unique characteristics which dictate their suitability for specific applications.

1.2 Lithium-Ion Batteries (LIBs)

Lithium-Ion Batteries (LIBs) represent the leading edge of current rechargeable battery technology, dominating the market for a vast range of applications due to their superior energy and power characteristics [16]. Their success stems from the high electrochemical potential of lithium and its ability to reversibly diffuse into various electrode materials [17-18].

1.2.1 Components of LIBs

A typical LIB cell consists of four main components as:

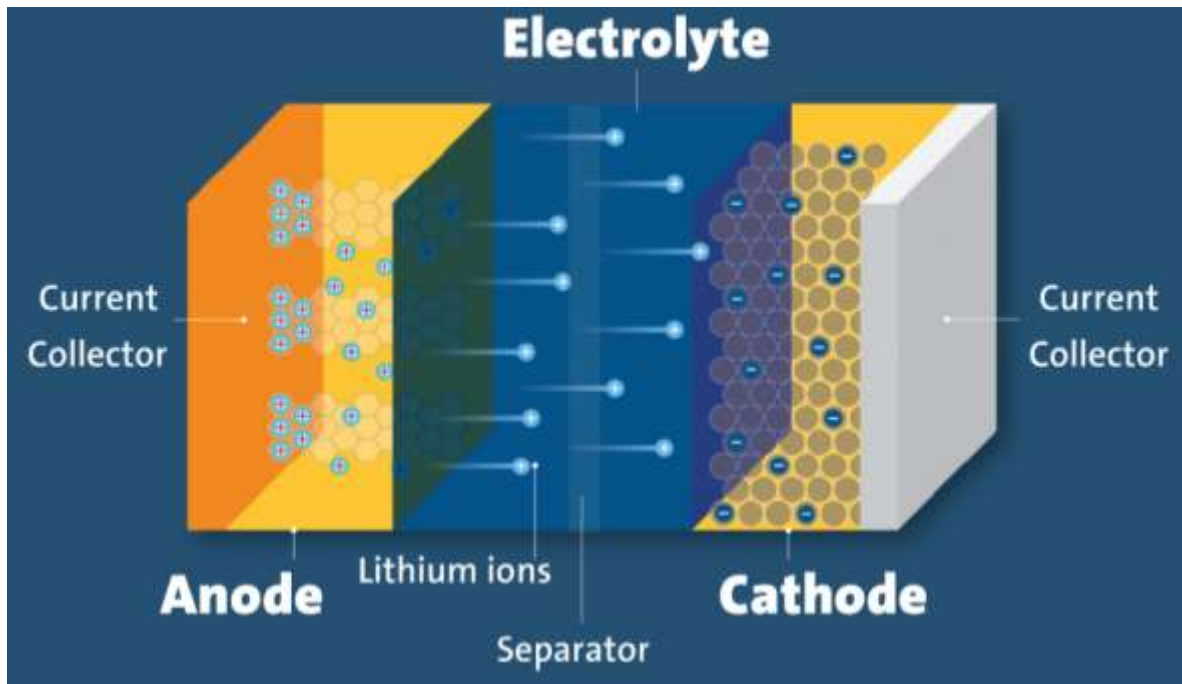


Fig. 1: Components of Lithium-Ion Battery

1.2.1.1 Electrolyte

The electrolyte facilitates the transport of lithium ions (Li^+) from anode to cathode and vice versa. It must possess high ionic conductivity and electronic insulation. The most common type of electrolyte in commercial LIBs is a non-aqueous solution consisting of lithium salts dissolved in organic solvents [19-20].

- **Lithium Salts:** The widely used lithium salt is lithium hexafluorophosphate (LiPF_6) because of its conductivity, stability, and ability for the formation of stable SEI. Other salts like lithium bis(trifluoromethanesulfonyl)imide (LiTFSI) and lithium bis(oxalato)borate (LiBOB) are also explored for specific properties like enhanced safety or improved high-temperature performance [21].
- **Organic Solvents:** Common organic solvents include cyclic carbonates and linear carbonates. Ethylene carbonate (EC) is crucial for forming a stable SEI layer, while Dimethyl carbonate (DMC) is often added to lower viscosity and improve ionic conductivity. The mixture of these solvents is carefully optimized to achieve a balance of properties [22].

- **Additives:** Small amounts of additives are often included to improve specific aspects of battery performance, such as enhancing SEI formation, improving high-voltage stability, or acting as flame retardants [23].

1.2.1.2 Cathode

The cathode (positive electrode) is typically a lithium-containing metal oxide. This material releases lithium ions when the battery discharges and restores them when it charges. Key requirements for cathode materials include high specific capacity, high operating voltage, good cycle stability, and thermal stability [24].

Common cathode materials include:

- **Lithium Cobalt Oxide (LiCoO_2 , LCO):** It is known for its high energy density and relatively simple synthesis. It is widely used in portable electronic devices but suffers from limited thermal stability and cycle life, especially at high charge rates, and contains expensive cobalt [25].
- **Lithium Manganese Oxide (LiMn_2O_4 , LMO):** A spinel structure material offering good thermal stability and lower cost. It has a high power capability but typically lower specific capacity compared to LCO and can suffer from manganese dissolution at elevated temperatures [26].
- **Lithium Nickel Manganese Cobalt Oxide (LiNiMnCoO_2 , NMC):** A ternary blend that combines the advantages of nickel (high capacity), manganese (stability, lower cost), and cobalt (stability, power). By varying the ratios of Ni, Mn, and Co, properties can be tuned for various applications. NMC is a dominant material in electric vehicles [27].
- **Lithium Iron Phosphate (LiFePO_4 , LFP):** Characterized by its exceptional thermal stability, long life, and low cost. Although its operating voltage is lower than other cathode

materials, leading to slightly lower energy density, its safety and longevity make it ideal for power tools, electric buses, and grid-scale energy storage [28].

Table 1: Analysis of cathode materials

Material	Theoretical Specific Capacity (mAh g ⁻¹)	References
LiCoO ₂	274	[25]
LiMn ₂ O ₄	148	[26]
LiNiMnCoO ₂	150-200	[27]
LiFePO ₄	170	[28]
LiMgMnO ₄	160-170	[29]
LiNiO ₂	275	[30]

1.2.1.3 Separator

The separator is a porous membrane that is placed between the anode and cathode for allowing the free flow of lithium ions through the electrolyte and to prevent electronic short-circuiting. It must be electronically insulating, chemically stable within the electrolyte, mechanically strong, and have good wettability with the electrolyte [31].

- **Materials:** Typically made from polymer materials such as polyethylene (PE) or multi-layered PE/PP structures. These materials are chosen for their chemical inertness and low cost.
- **Pore Structure:** The separator has a precise pore structure that allows for optimal ion transport and to prevent short-circuiting.

1.2.1.4 Anode

The anode (negative electrode) is the host material for lithium ions. During charging, commonly known as intercalation these ions are stored in the anode material and released during discharging, commonly known as de-intercalation. The ideal anode material should have high specific capacity, good cycle stability, low volume expansion upon lithiation, and a safe operating potential close to that of metallic lithium.

- **Graphite:** The most common anode material is Graphite is favoured due to its excellent electrical conductivity, good cycling performance, relatively low cost, and relatively low volume change during lithium intercalation. Lithium ions insert into the layered structure of graphite to form LiC_6 [32].
- **Silicon (Si):** Silicon has a theoretical specific capacity significantly higher than graphite. This high capacity makes silicon a promising next-generation anode material. However, it suffers from large volume changes, leading to mechanical degradation and rapid capacity fade. Research focuses on silicon or combining it with graphite to mitigate these issues [33].
- **Lithium Titanate ($\text{Li}_4\text{Ti}_5\text{O}_{12}$, LTO):** A spinel oxide material known for its exceptional cycle life because of its "zero-strain" intercalation mechanism. However, its higher operating potential leads to a lower cell voltage and thus lower energy density compared to graphite anodes. LTO is used in applications requiring long life and high safety, such as electric buses and energy storage systems [34].

Table 2: Analysis of different anode materials

Material	Theoretical Specific Capacity (mAh g ⁻¹)	References
Graphite	960	[32]
Silicon	4200	[33]
Porous Carbon	800-1000	[35]
Carbon Nanotubes	1100	[36]
SiO	1600	[37]
Transition Metal Oxides	500-1200	[38]

1.3 Principle and Working

The working principle of a LIB relies on the reversible diffusion of lithium ions into and out of the electrode materials. This process does not involve the formation of metallic lithium, hence the term “Lithium-Ion Battery” [39].

1.3.1 Electrochemistry

During discharge (when the battery provides power), ions de-intercalate from the anode and intercalate into the cathode material by passing through the electrolyte and separator. Simultaneously, electrons are released from the anode, flow through the external circuit, and re-enter the cathode to balance the charge [40].

Anode reaction for Graphite:



Cathode reaction for LCO:

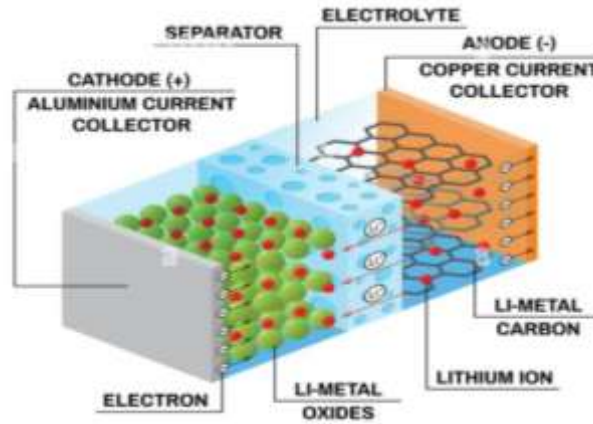
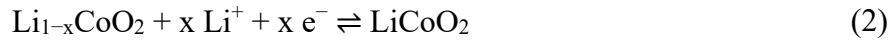


Fig. 2: Dis-charging process in LIBs

During charging, the external power source reverses the process. The ions de-intercalate from the cathode, migrate back and intercalate into the anode material. Electrons flow from the cathode through the external circuit to the anode [41].

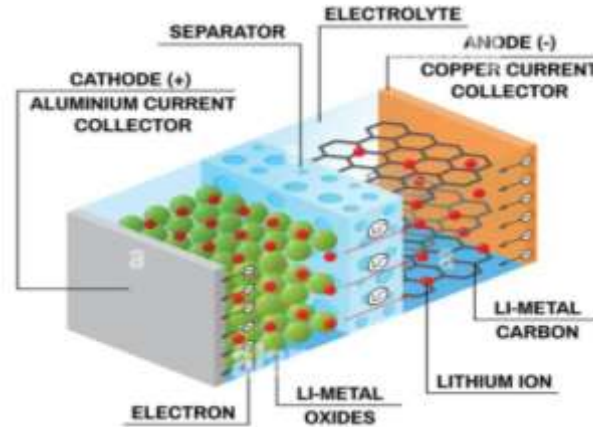
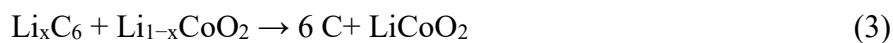


Fig. 3: Charging process in LIBs

For typical LIBs, the average operating voltage is around 3.7 V [42]. The overall electrochemical reaction for a LIB cell can be represented as:

Overall Reaction (Discharge):



1.3.2 Electrochemical Performance

The performance of a LIB is characterized by several key metrics [43]:

1.3.2.1 Charge/Dis-charge

- **Charge Cycle:** The process of supplying electrical energy to the battery to restore its stored chemical energy.
- **Discharge Cycle:** The process of the battery releasing its stored chemical energy as electrical energy.
- **C-rate:** The rate at which a battery is charged or discharged with relation to its maximum capacity of the active material. Lower C-rates gives the structural stability of the electrode material whereas higher C-rates generally lead to reduced effective capacity.

1.3.2.2 Specific Capacitance

Specific capacitance refers to the capability of a material to store electrical charge in relation to its unit mass, usually expressed in Farads per gm ($F\ g^{-1}$).

Specific capacity is the measurement of the maximum amount of electrical charge that can be stored in relation to its unit mass, stated as milliampere-hours per gram ($mAh\ g^{-1}$). It is a crucial parameter for determining the energy density of a battery [44].

1.3.2.3 Lithium-Ion Diffusion Coefficient

It quantifies the rate of diffusion of lithium ions to move through electrode-electrolyte interface. A higher diffusion coefficient indicates faster ion transport, which is critical for achieving high power density and efficient charge/discharge at high C-rates [45]. The kinetics of lithium ion transport through solid electrodes is taken as the rate-determining step. Various electrochemical techniques, such as Electrochemical Impedance Spectroscopy (EIS), Cyclic Voltammetry (CV) and Galvanostatic Intermittent Titration Technique (GITT), are used to determine the diffusion coefficients [46-47].

Chapter 2

LITERATURE REVIEW

This chapter gives a conceptual idea of existing literature pertinent for characterization of electrode materials for energy storage applications. A particular focus is placed on various types of anode materials, their electrochemical mechanisms, and their performance characteristics.

2.1 Types of Anodes

Anode materials are critical components in LIBs for enhanced electrochemical performance. The selection of an anode material is crucial as it significantly dictates the overall energy density and long-life cycle of the device. This section reviews prominent categories of anode materials [48].

Intercalation-Type Anodes

Intercalation-type anodes store lithium ions reversibly within their layered structures without undergoing significant structural changes or forming new chemical bonds with the host material. The lithium ions insert themselves into vacant sites or interstitial spaces within the host lattice in the charging and subsequently extracted out in the discharging process. This "rocking-chair" mechanism, where lithium ions shuttle between the anode and cathode, relies on the structural integrity of the host material [49]. For example, Graphite.

The general reaction for an intercalation anode (M) is represented as:



During lithiation, lithium ions and electrons are simultaneously inserted into the host lattice. During delithiation, they are extracted. The process is typically reversible, leading to good cycling stability.

Alloying-Type Anodes

Alloying-type anodes store lithium ions through a reversible alloying reaction, where lithium combines with the electrode material, forming a distinct lithium-rich alloy phase. This mechanism typically involves a high number of lithium ions per host atom, leading to significantly higher theoretical specific capacities than intercalation-type materials [50]. For example, Silicon.

The general reaction for an alloying anode (M) represented as:

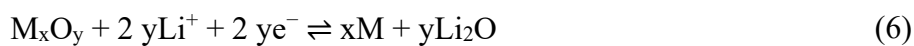


During lithiation, the host material alloys with lithium to form a $Li_x M$ phase, undergoing a phase transformation. During de-lithiation, the alloy phase de-alloys, releasing lithium ions and typically returning to the original host material or a stable intermediate phase.

Conversion-Type Anodes

Conversion-type anodes store lithium ions via a reversible conversion reaction, where the host material (typically a metal oxide, sulfide, phosphide, or nitride) is fully decomposed into metallic nanoparticles and an amorphous lithium compound (e.g., Li_2O , Li_2S , Li_3P). This mechanism involves the breaking of existing bonds and the formation of new ones, typically involving multi-electron transfer, which contributes to very high theoretical capacities [51]. For example, Transition Metal Oxides (TMOs).

The general conversion reaction for a metal oxide (M_xO_y) in a LIBs can be represented as:



During lithiation, the metal oxide is reduced to finely dispersed metallic nanoparticles (M) embedded within an amorphous Li_2O matrix. Upon de-lithiation, the reverse oxidation reaction occurs, ideally regenerating the original metal oxide.

2.1.1 Transition Metal Oxides (TMOs)

TMOs have garnered significant attention as anode materials. Unlike traditional graphite anodes that primarily rely on intercalation mechanisms, many TMOs operate via a conversion reaction mechanism. This involves the reversible breaking and reforming of metal-oxygen bonds during the charge and discharge processes [52].

Table 3: Common TMO anodes along with their advantages and di-advantages

TMOs	Advantages	Dis-advantages	Ref.
Cobalt Oxide (Co ₃ O ₄ , CoO)	high theoretical capacities	volume changes during cycling and capacity fade	[53]
Nickel Oxide (NiO)	high theoretical capacities	large volume expansion and poor electrical conductivity	[54]
Iron Oxide (Fe ₂ O ₃ , Fe ₃ O ₄)	low cost and economically cheap	moderate cycling stability due to volume changes	[55]
Manganese Oxide (MnO ₂ , Mn ₃ O ₄)	high theoretical capacities	intrinsic poor electrical conductivity and rate capability	[56]

Despite their high theoretical capacities, practical application of TMOs as anodes faces several challenges:

- **Large Volume Change:** The conversion reaction often involves substantial volume expansion and contraction, leading to mechanical degradation.
- **Poor Electrical Conductivity:** Most TMOs are semiconductors or insulators, resulting in high internal resistance and limited rate capability.
- **Solid Electrolyte Interphase (SEI):** Repeated volume changes can continuously break and reform the SEI layer, consuming active lithium and leading to irreversible capacity loss.

2.1.2 Mixed Transition Metal Oxides (MTMOs)

Mixed transition metal oxides (MTMOs), also known as complex or composite transition metal oxides, are materials composed of two or more different transition metal cations in a single oxide lattice. [57] These materials are emerging as promising anode material because of their ability to combine the advantages of individual TMOs while often mitigating their respective drawbacks. The presence of multiple metal centres can lead to enhanced electrochemical activity, improved structural stability, and superior electronic conductivity as compared with single TMOs. The synergistic effect happening between the different metal ions contribute to their improved electrochemical performance.

MTMOs primarily operate by a conversion reaction mechanism to single TMOs, but the involvement of multiple metal species adds complexity and offers potential for superior performance [58-59].

The presence of multiple metal cations can lead to enhanced lithium storage sites, Improved electrical conductivity, and better structural integrity.

Among them, compounds with the general formula AB_2O_4 , exhibiting a spinel-structured within $Fd3m$ space group, have gained attention [60]. In these materials, A cation occupies tetrahedral sites and B cation occupies octahedrally sites, both coordinated by oxygen atoms, allowing efficient lithium-ion diffusion pathways. The coexistence of two distinct transition metals in the single-phase crystal structure enables multiple redox couples, thereby improving the specific capacity, rate capability, and cyclic stability of the anode material [61-62].

2.1.2.1 Nickel Manganese Oxide ($NiMn_2O_4$)

Among the various MTMOs, a spinel AB_2O_4 -type material, NMO ($NiMn_2O_4$) is widely used as an active anode material. This is primarily due to the advantageous combination of nickel's high theoretical capacity and manganese's lower cost, environmental benignity, and relatively stable

redox chemistry. The mixed valence states of Ni and Mn, along with their different ionic radii, contribute to a complex structure that can potentially offer improved performance [63-64].

Structure

A spinel AB_2O_4 -type structure of NMO where Ni cation occupies tetrahedral sites and Mn cation occupies octahedral sites, both coordinated by oxygen atoms.

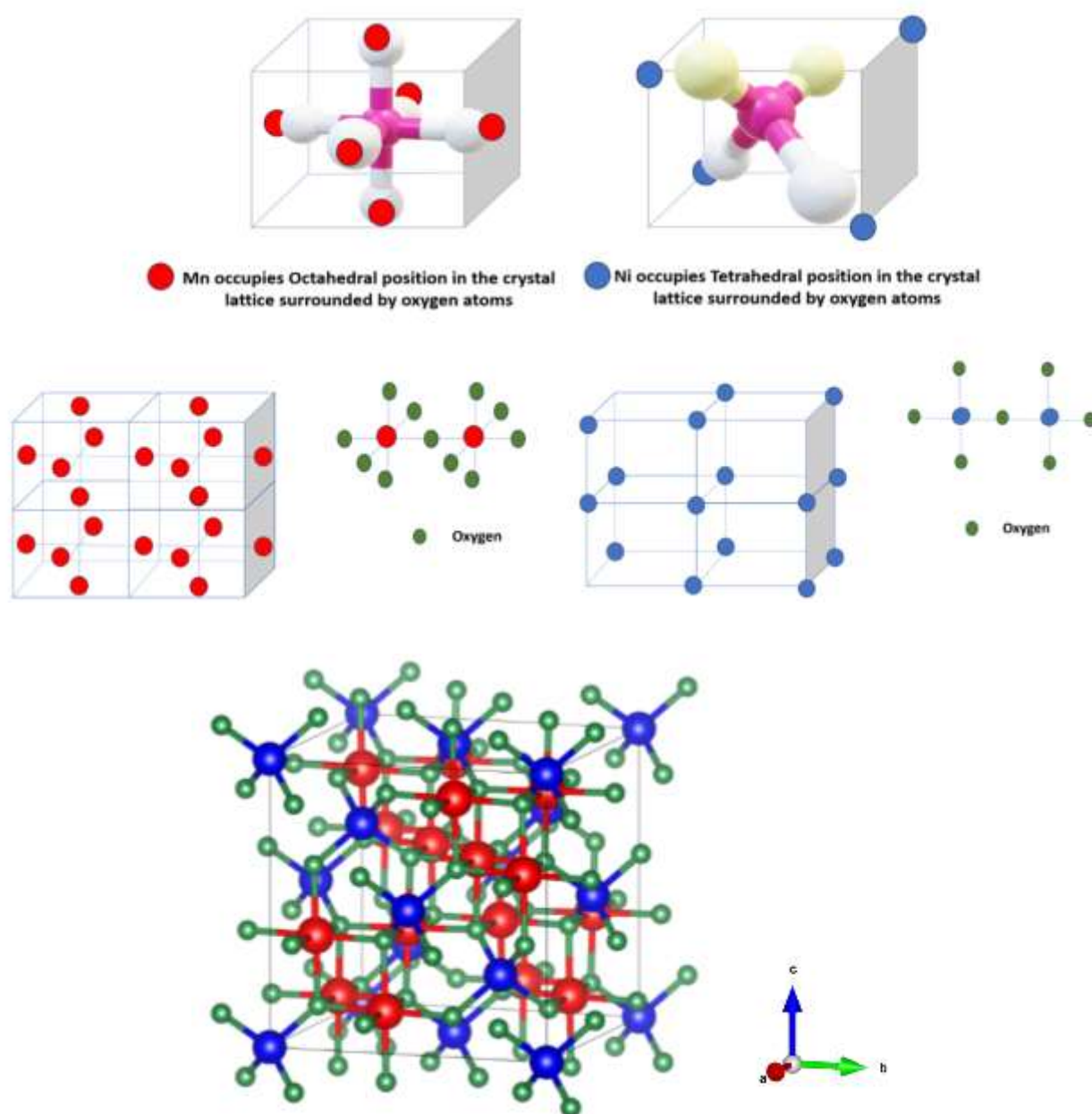
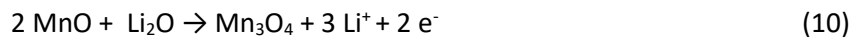
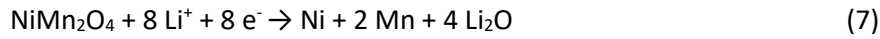


Fig. 3: Structural analysis of NMO

Mechanism

In a typical reaction for NiMn_2O_4 :



During lithiation, both Ni and Mn ions are reduced, forming metallic Ni and Mn particles in a Li_2O matrix. During de-lithiation, the reverse oxidation of Ni and Mn occurs. The variable oxidation states of Ni and Mn, and their interplay contribute to the overall charge storage capacity. The manganese help to stabilize the structure during repeated volume changes, while nickel contributes significantly to the capacity. However, this process can also lead to structural instability, as the material undergoes changes in phase and oxidation and oxidation state [65]. These transformations can induce volume changes in the electrode, which may contribute to capacity fading and reduced cycling stability over time. On the other hand, insertion process occurs within the voids and interstitial spaces of the cubic NMO structure, where lithium ions are reversibly inserted and extracted [66-67]. This process is typically more structurally stable, though it may offer lower capacity than conversion reactions.

Both the processes contribute significantly to the overall performance of the NMO electrode during cycling. An important factor influencing these processes is the lithium-ion diffusion, which quantifies the movement or flow of lithium ions within the electrode. This coefficient is crucial for understanding the efficiency, conductivity and performance of the material during cycling and is calculated through electrochemical analysis [68-70].

Chapter 3

EXPERIMENTAL PROCEDURE

This chapter gives an outline the materials utilized for synthesizing NiMn_2O_4 as anode material and fabricating Lithium-half cell for LIBs. It covers all the structural, morphological, and electrochemical analysis for the characterization of the electrode material and fabricated electrode.

3.1 Materials and Methods

Nickel (II) acetate tetrahydrate ($\geq 99\%$) and Manganese (II) acetate tetrahydrate ($\geq 99\%$) were brought from Merck, India. Citric Acid ($\geq 99\%$) was purchased from Thermo–Fischer Scientific, India. All the chemicals used without purification.

The precursors of Nickel and Manganese, along with Citric acid were weighed properly. A stoichiometric amount of Ni and Mn precursors were used in the ratio of 1:2. The ratio of chelating agent was kept 1:1 with the total metal ions. The simple hand grinding (solid-state) method was used to synthesis NMO. All the precursors along with chelating agent were hand grinded using mortar-pestle for around 8 hours. After grinding the sample, it was calcinated for 3 hours in a furnace at 800°C .



Fig. 5: Schematic representation of synthesis process

3.2 Material Characterization

The crystallinity and phase of the sample was analysed by XRD. The crystallite size of synthesised NMO was measured by using Scherrer equation and Williamson-Hall (W-H) plot approach, further analysed by SAED. The morphology and average particle size of the synthesised NMO was studied by SEM and TEM, and elemental mapping of synthesized sample were determined by EDS.

3.2.1 X-ray Diffraction (XRD)

XRD is an analytical technique used to identify the structural characteristics of the materials. The principle is based on the constructive interference of x-rays (x-ray of single wavelength) with a sample. When these incident x-rays strike a crystal lattice, they get diffracted at specific angles that are calculated as per Bragg's Law:

$$n\lambda = 2d\sin\theta \quad (11)$$

In a typical setup, x-rays are generated from commonly Cu $k\alpha$ radiation, and then directed towards the crystalline sample. The sample is mounted on a goniometer, which precisely controls the angle of incidence (θ) and the angle of detection (2θ). As the sample rotates, different crystallographic planes satisfy Bragg's Law, leading to diffracted beams of varying intensities. XRD provides data on the phase identification of crystalline materials, lattice parameters, and crystallite size.

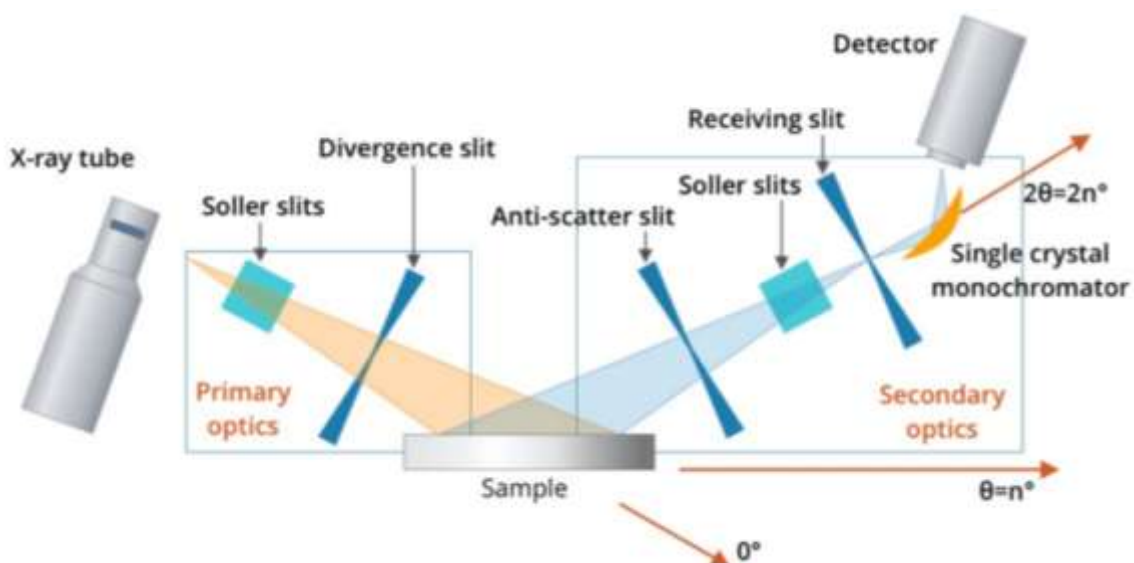


Fig 6: Diagrammatic representation of XRD

3.2.2 Scanning Electron Microscopy (SEM)

SEM is a technique that produces high-resolution images at micro- to nano-meter scale of the surface of the sample, by bombarding it with a finely focused beam of electrons. When the incident electron beam interacts with the sample, various signals are produced, including secondary electrons (SEs), backscattered electrons (BSEs), and characteristic X-rays. SEM primarily utilizes SEs and BSEs for imaging. SEs are low-energy electrons emitted from the very surface of the sample, providing topographical information. BSEs are high-energy electrons from the incident beam that are scattered back from the sample elastically, providing compositional information based on atomic number contrast. It is especially useful for studying surface features such as grain boundaries, particle size, and texture.

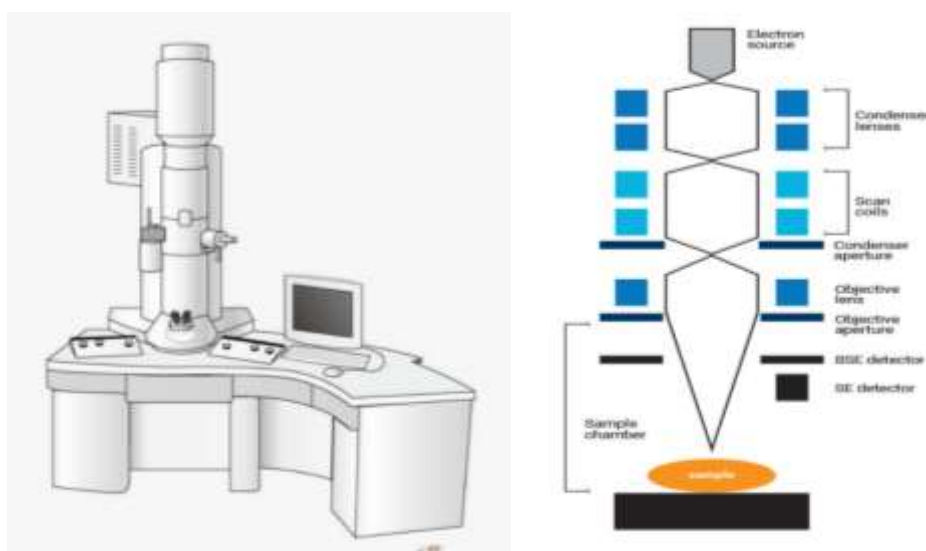


Fig. 7: Diagrammatic representation of SEMs

3.2.3 Energy Dispersive X-ray Spectroscopy (EDS)

EDS is an analytical technique used for elemental analysis of a sample. It is typically coupled with SEM. Its principle is based on the generation of characteristic X-rays when the incident electron beam interacts with the atoms in the sample. When a high-energy electron beam strikes an atom, it can eject an inner-shell electron, creating a vacancy. An electron from a higher energy shell then drops into this

vacancy to fill it, releasing the excess energy as a characteristic X-ray photon, whose energy is unique to the element from which it originated. EDS provides qualitative and semi-quantitative elemental analysis of the sample. It helps in determining the elemental composition of elements within a sample.

3.2.4 Transmission Electron Microscopy (TEM)

TEM is a powerful characterization technique that uses a beam of electrons transmitted through an ultra-thin sample to obtain high-resolution images, diffraction patterns, and elemental information. Unlike SEM, which images the surface, TEM provides internal structural and morphological information. The principle relies on the wave-particle duality of electrons, a high-energy electron beam (with a very short wavelength) can be transmitted through a sample, and the interaction of these electrons with the sample structure provides information about its internal morphology, crystal structure, and defects. TEM provides detailed internal structural information at atomic to nanometer resolution, including crystal structure, lattice defects, grain boundaries, and morphology. It is suitable for both imaging and diffraction studies.

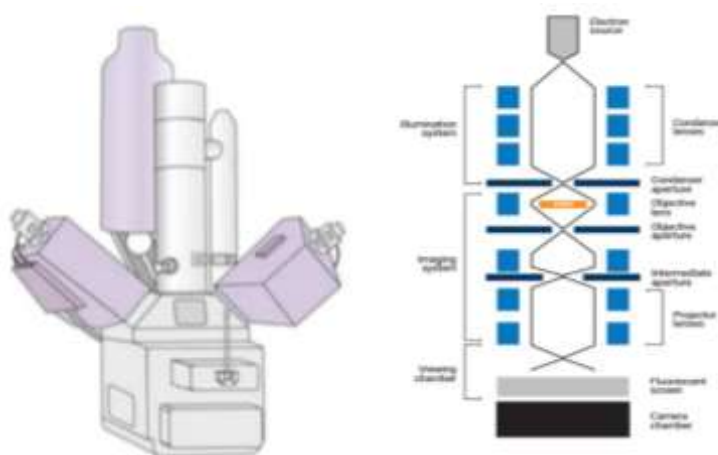


Fig. 8: Diagrammatic Representation of TEM

3.2.5 Selected Area Electron Diffraction (SAED)

SAED is a technique performed within a TEM to obtain crystallographic information from a specific, small region of a thin sample. Its basic principle is based on the diffraction of electrons by the periodic atomic planes within a crystalline material. Similar to XRD, constructive interference occurs when

electrons are scattered by atomic planes. However, due to the much shorter wavelength of electrons compared to X-rays, the diffraction angles are much smaller, and the diffraction pattern consists of spots or rings rather than continuous peaks. SAED provides insights on the crystal structure and phase identification of materials. It is particularly useful in identifying crystalline phases and analysing polycrystalline or nanocrystalline samples.

3.3 Electrode Fabrication

The working electrode was fabricated using composite of Active material (AM), Activated Carbon (AC), along with Polyvinylidene Fluoride (PVDF) as binder in a ratio of 8:1:1 ratio. The components were thoroughly mixed in N-methyl Pyrrolidone (NMP) solvent to form a homogeneous slurry. This slurry was uniformly coated on Cu foil (current collector) using doctor blade and dried at 120°C for 12 hours to ensure complete evaporation of the solvent and proper adhesion of the electrode material. After drying, circular electrodes of 16 mm diameter were punched out from the coated foil. These electrodes were then assembled into coin cells inside an argon-filled glove box workstation (MBraun) with O₂ and H₂O level below 0.5 ppm to prevent contamination. A coin cell of standard CR2016 configuration with 20 diameter and 1.6 mm thickness was used for assembling. A lithium metal chip was used as both the reference and counter electrode. A Celgard 2400 separator (Polypropylene membrane) was placed between working and lithium chip to prevent short-circuiting while allowing ionic conductivity. The electrolyte employed in the cells was 1M LiPF₆ in organic solvent EC:DMC in 1:1, by volume ratio. Additionally, a metal spacer was inserted to accommodate any excess internal space and ensure firm contact between components.

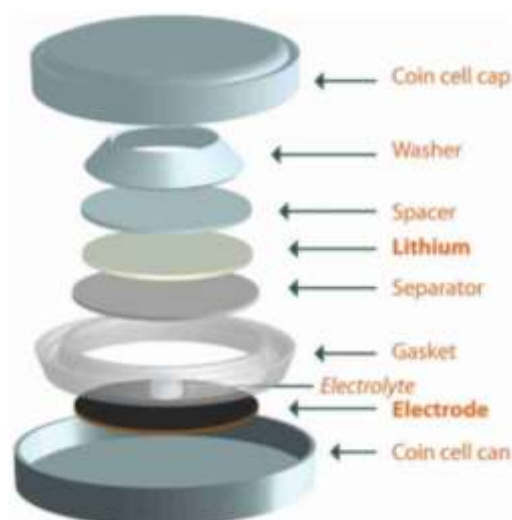


Fig. 9: Coin-cell assembly representation

3.4 Electrochemical Measurement

Electrochemical techniques are a category of analytical methods that utilize electrochemical principles to analyze and examine chemical reactions, investigate substance's properties, and determine concentrations. Electrochemical techniques present more benefits than technologies of surface modification. The electrochemical methods consist of measurements, Electrochemical Impedance Spectroscopy (EIS), Cyclic Voltammetry (CV), and Galvanostatic Charge/Dis-charge. In this work, the EIS and CV measurements were performed by BioLogic Potentiostat and software EC Lab. The GCD characteristics were evaluated by Lithium-Ion Battery Cycler system.

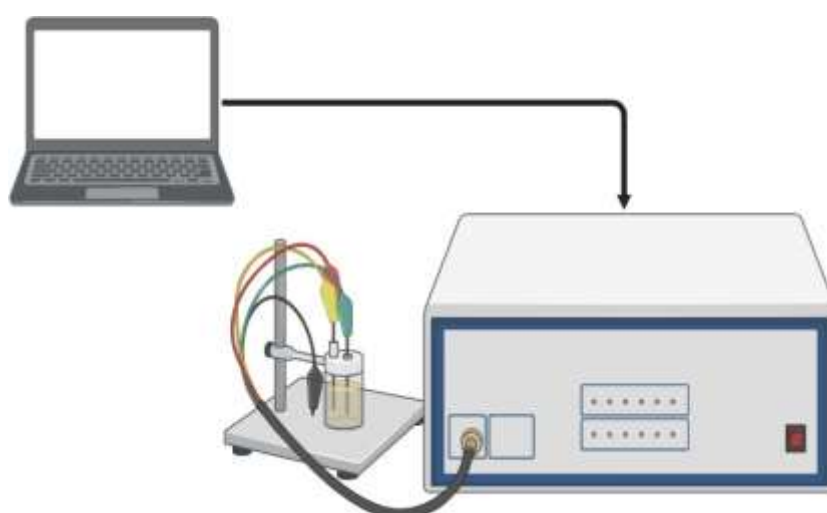


Fig. 10: Diagrammatic Representation of Potentiostat for Electrochemical analysis

3.4.1 Electrochemical Impedance Spectroscopy (EIS)

EIS is a technique employed to investigate the electrochemical properties of material by using a small AC voltage and measuring the resultant AC current over low and high frequencies. The fundamental principle involves disturbing an system from its equilibrium state with a sinusoidal potential and analysing the response as a function of frequency. The impedance gives information about the processes occurring at the electrode-electrolyte interface, including charge transfer kinetics, solution resistance, and lithium-ion diffusion coefficient. EIS data are primarily visualized through Nyquist and Bode plots, which are subsequently analysed using equivalent circuit modelling.

3.4.2 Cyclic Voltammetry (CV)

CV analysis investigates the redox behaviour and reaction kinetics of electroactive species by sweeping the potential of the working electrode linearly within potential window at a constant scan rate, while simultaneously measuring the resultant current. The principle involves employing a potential waveform to the electrode, causing the potential to sweep in one direction (anodic scan) and then reverse and sweep back in the reverse direction (cathodic scan). As the potential is swept, electroactive species at the electrode surface undergo oxidation or reduction, leading to characteristic current responses. The shape and position of the resulting current-potential peaks provide information about reaction reversibility, reaction mechanisms, and diffusion processes. The area under the CV curve can be correlated with the specific capacitance of the electrode material.

3.4.3 Galvanostatic Charge/Dis-charge (GCD)

GCD is an electrochemical method employed to evaluate the capacity and cycling stability of energy storage devices. The principle involves applying a constant current to the cell and monitoring the resultant potential response as a function of time. During charge, a constant positive current is applied, resulting in potential increases, while during discharge, a constant negative current is applied, and the resulting potential decreases. The shape of the voltage profile gives information about the charge storage mechanism of the device.

RESULTS AND DISCUSSIONS

4.1 Structural Analysis

The NMO phase was confirmed through XRD analysis as shown in Fig.11, revealing the cubic lattice and $Fd3m$ space group, perfectly matching with JCPDS 01-071-0852. The reflections (111), (220), (311), (222), (400), (422), (511), (440), (531), (622), (444), and (731) corresponding to NMO. The sharp and well-defined peaks indicate that the synthesised material possesses high crystallinity and it may result with enhanced structural stability during electrochemical reactions. The W-H plot and Scherrer's equation both were employed to calculate the crystallite size of the synthesised sample (Fig. 12). The W-H plot gives a negative slope indicating that there is a lattice compression within the material.

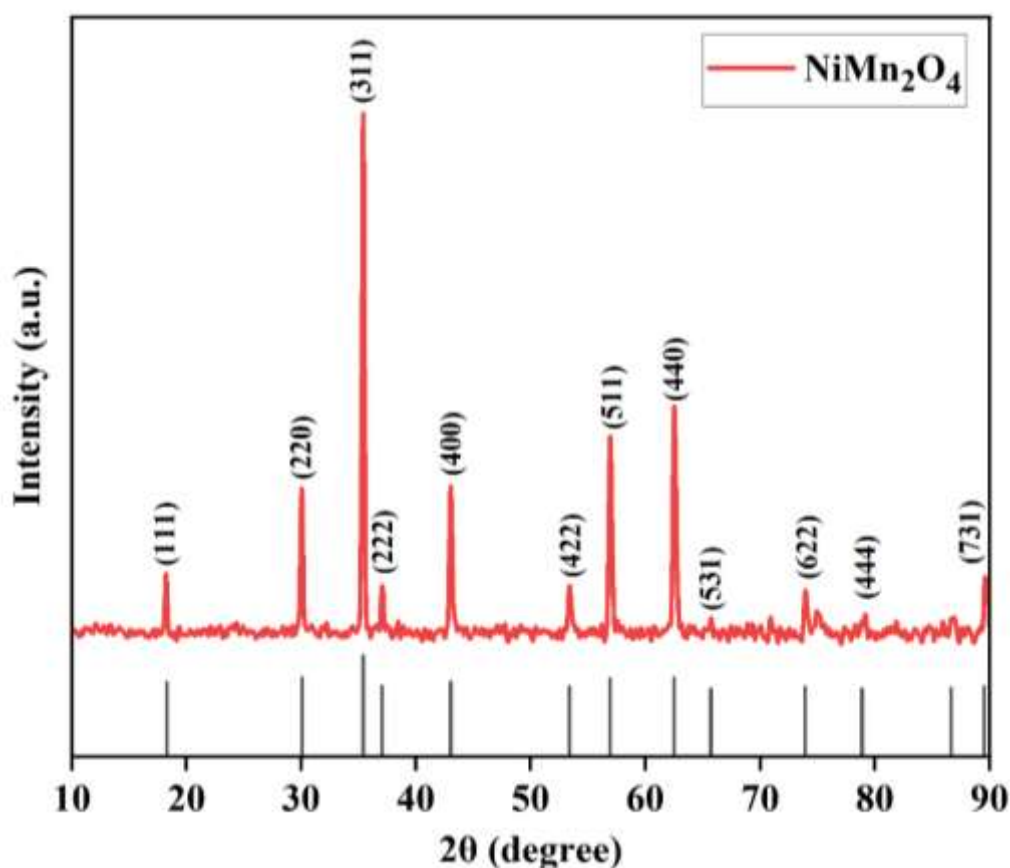


Fig. 11: X-ray analysis of NMO calcinated at 800°C for 3 hours.

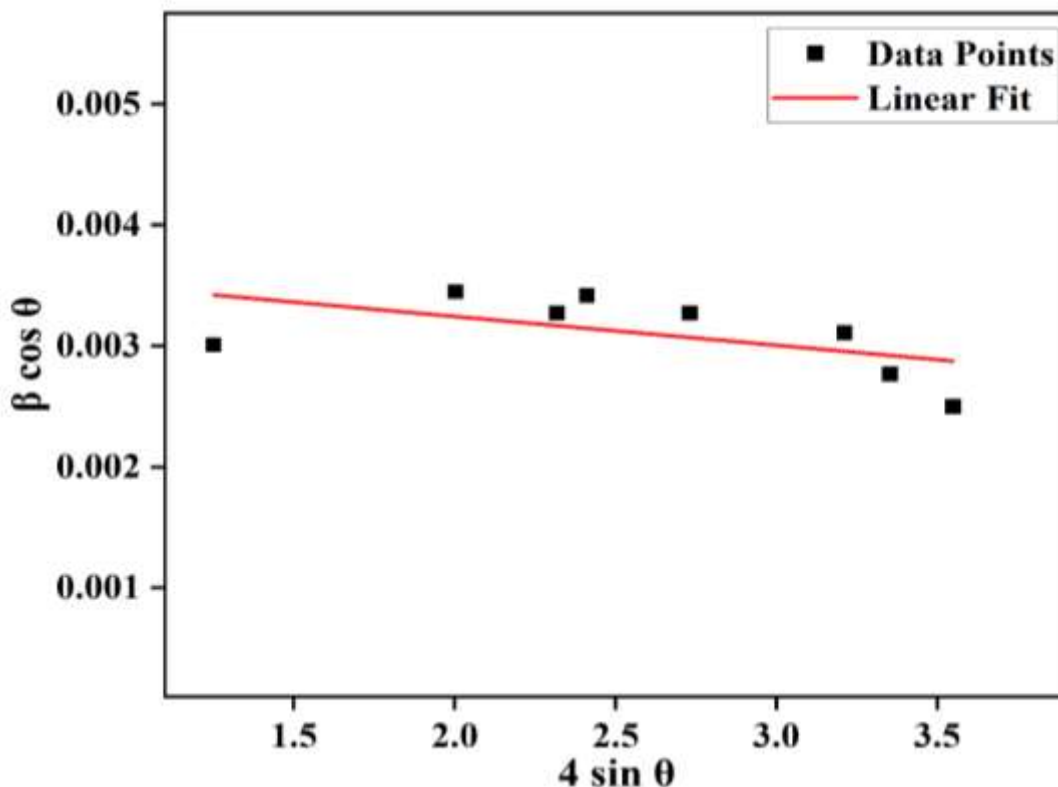


Fig. 12: Williamson-Hall (W-H) plot of synthesised sample

4.2 Morphological Analysis

The morphology and particle size of the synthesised NMO were thoroughly investigated using SEM and TEM analysis. The SEM images, presented in Figure 13, reveal the overall architecture, particle size, and surface characteristics of the synthesized NMO material. As shown in Fig. 13 (a-b), the material exhibits a relatively uniform morphology, primarily composed of hexagonal shape. The average particle size is estimated to be approximately 282.32 nm, with a size range observed between 250 to 350 nm. At higher magnification, as seen in Fig. 13 c), the individual particle surface appears as a proper hexagon shaped, and distinct inter-particle voids are visible, contributing to the overall porosity of the material. This observed morphology is crucial for indicating the material's electrochemical performance. The extent of porosity and voids facilitates the diffusion of lithium ions through the electrode material, thereby reducing diffusion pathways and potentially enhancing the material's rate capability. The overall architecture appears to provide sufficient connectivity for efficient electron transport within the electrode.

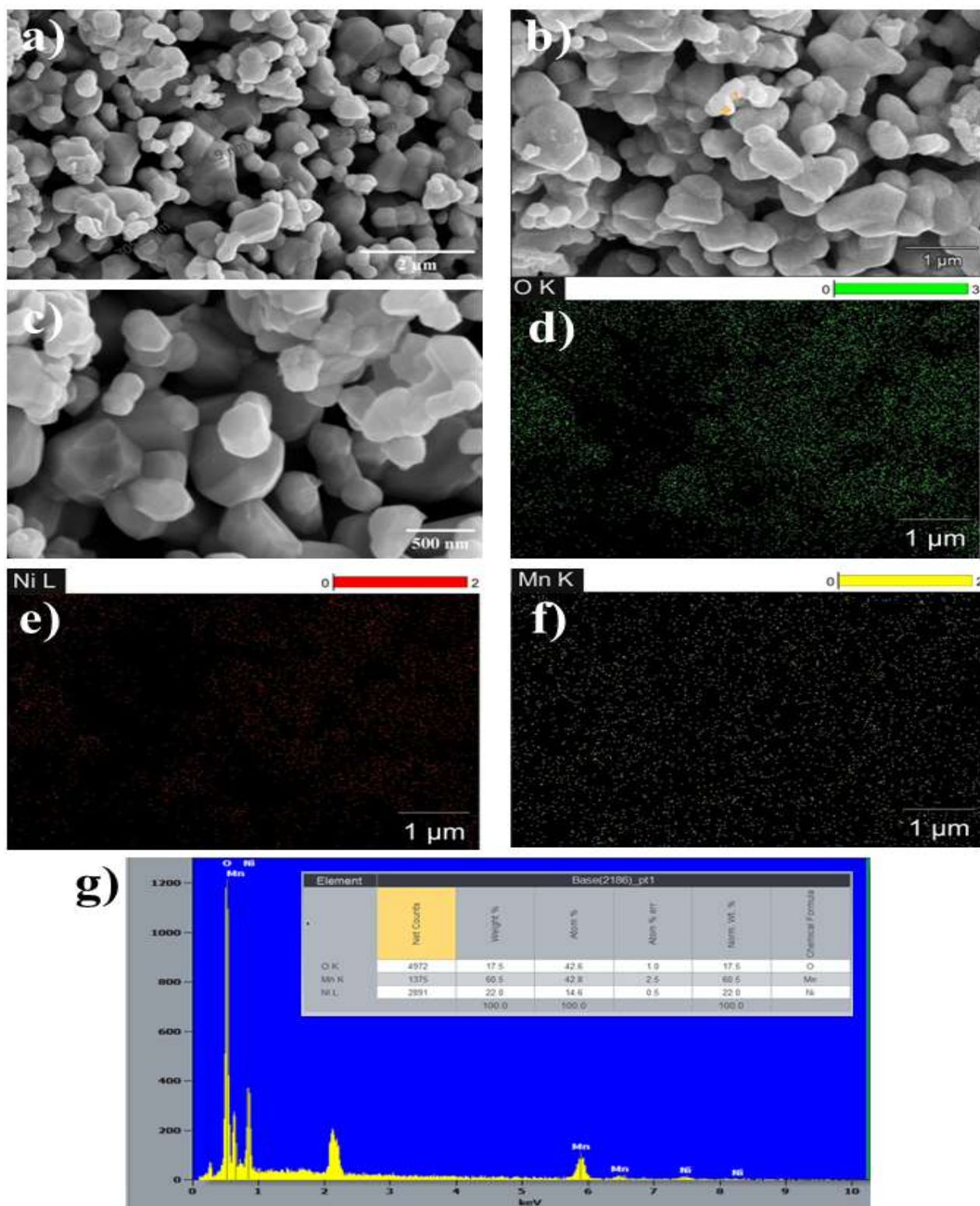


Fig 13: SEM analysis of synthesised NMO at a) 50k x, b) 100k x, c) 120k x, and elemental mapping images of d) Oxygen (green), e) Nickel (red), f) Manganese (yellow) and g) EDS analysis along with map sum spectrum

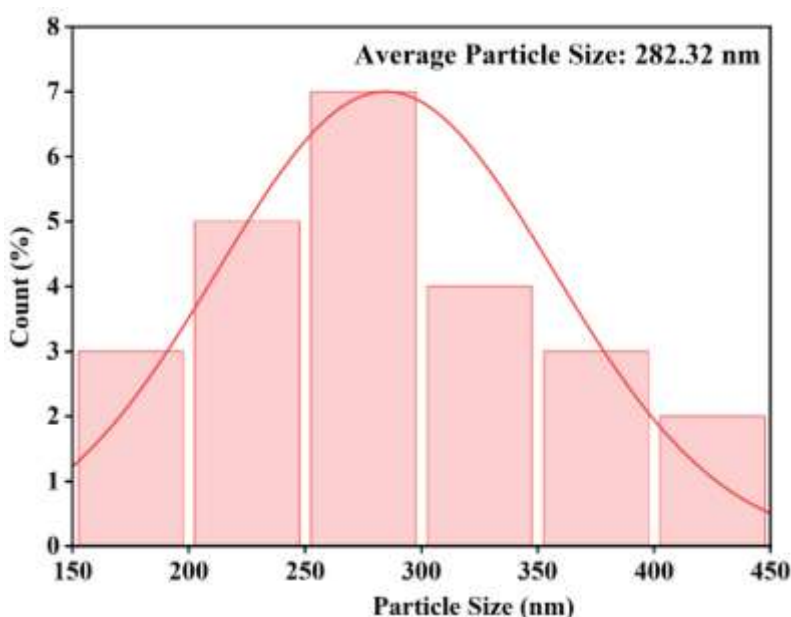


Fig 14: SEM histogram for average particle size

TEM analysis (Fig. 15 (a,b)) shows the TEM results of NMO particles at different magnification along with SAED pattern. The images confirmed the hexagonal morphology and revealed a smaller average particle size of 180.52 nm. The crystalline structure of the NMO particles was investigated using both XRD and SAED pattern respectively. The d-spacing calculated from the SAED pattern were found to be in excellent agreement with the d-spacings obtained from the XRD data with matching the lattice planes. This strong correlation confirms that the NMO possess the expected crystalline phase. Despite confirming the correct phase and consistent d-spacing.

Morphological analysis conducted using SEM revealed that the as-synthesized nanoparticles have an average particle size of approximately 282.32 nm. However, TEM analysis, which involved a more vigorous dispersion step, showed a smaller average particle size of 186 nm. This discrepancy is commonly observed and suggests that while the primary particles are indeed around 180.52 nm, they exhibit a tendency to form loosely bound agglomerates or aggregates in their dry state, which are captured by SEM at the larger 282 nm scale. SEM, TEM analysis describes uniform microscale particles that directly supports the lithium-ion diffusion pathways and increased electrode-electrolyte contact area, thereby positively influencing the material's electrochemical performance. The

morphological analysis has been added in the revised manuscript with “RED” text as separate subsection in results and discussions. The observed crystalline size and particle size along with observed porosity directly supports improved lithium-ion diffusion pathways and increased electrode-electrolyte contact are, thereby influencing charge storage capability of NMO as an anode.

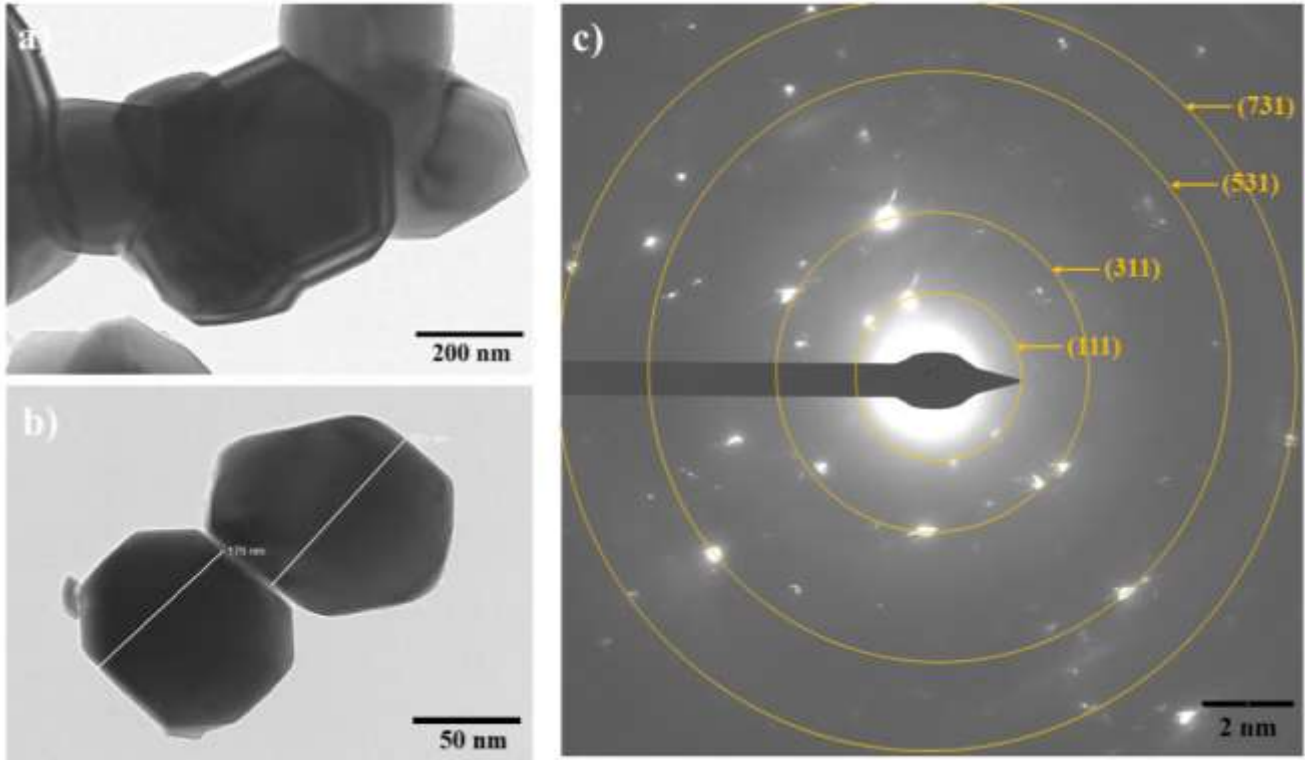


Fig 15: TEM analysis of NMO at a) 200nm, b) 50 nm, and c) SAED analysis at 2 nm range

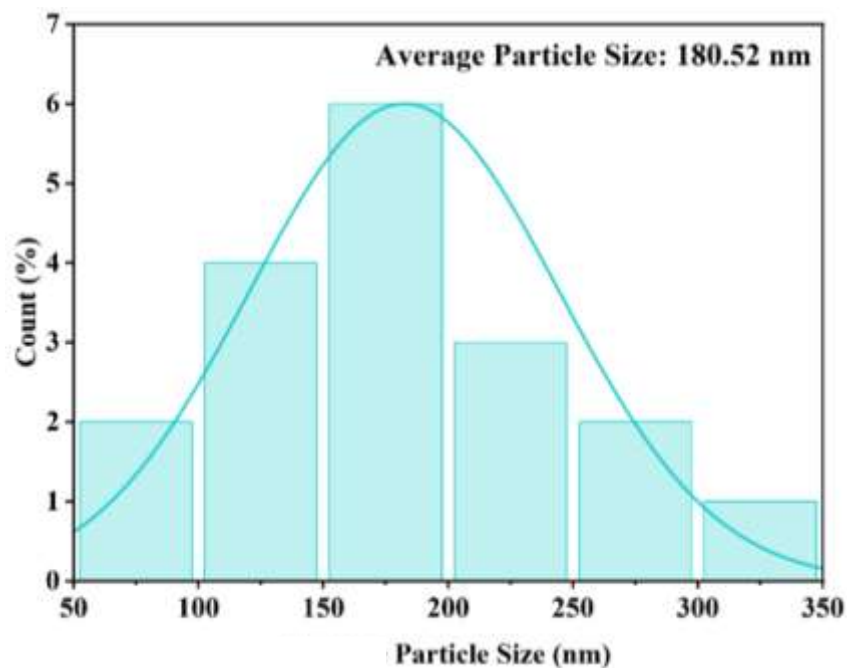


Fig 16: TEM histogram for average particle size

4.3 Electrochemical Analysis

The lithium-ion diffusion coefficient is calculated through two techniques, EIS and CV. Since, both the techniques are used to calculate Li^+ diffusion coefficient but their underlying principles differs. EIS uses an AC voltage to probe impedance across a range of frequencies. In the low frequency region, diffusion appears as a Warburg impedance, characterized by a linear region in the Nyquist plot. Whereas, the CV involves sweeping a voltage and measuring corresponding current. The peak currents in the anodic and cathodic region are diffusion-limited and are directly related to Li^+ diffusion coefficient. Hence, both of these techniques are employed in this research work to study the Li^+ diffusion coefficient of the synthesised material.

4.3.1 EIS analysis

EIS tests are shown in fig. 17, measured at 5mV AC amplitude, between the frequency interval of 1 MHz to 100 Hz. To measure the ionic conductivity, following relation (12) has been used:

$$\sigma = \frac{t}{R \cdot A} \quad (12)$$

where, σ , R , t , and A represents conductivity, bulk resistance, thickness and surface area of electrode, respectively.

The circuit fitting was done using EIS spectrum analyser software. In the Nyquist plot, it is observed that in the high-frequency region, there is a semi-circle showing resistance and in low- frequency domain, a straight line corresponding to the Warburg impedance. The Warburg Coefficient (σ_w) was obtained from the slope of the real part of the impedance (Z') versus the inverse square root of frequency in the low-frequency region. The Li^+ diffusion coefficient is evaluated by Warburg coefficient with the help of following equation (13):

$$D_L = \left(\frac{0.5 R^2 T^2}{n^2 \cdot F^2 \cdot A^2 \cdot \sigma_w^2 \cdot C_L^2} \right) \quad (13)$$

where, D_L , R , T , F , n , A , σ_w , and C_L represents lithium-ion diffusion coefficient, gas constant, absolute temperature (in K), Faraday constant, no. of electrons involved, surface area of the electrode, Warburg coefficient, and concentration of lithium, respectively. The lithium-ion diffusion was observed to be around $10^{-14} \text{ cm}^2 \text{ s}^{-1}$.

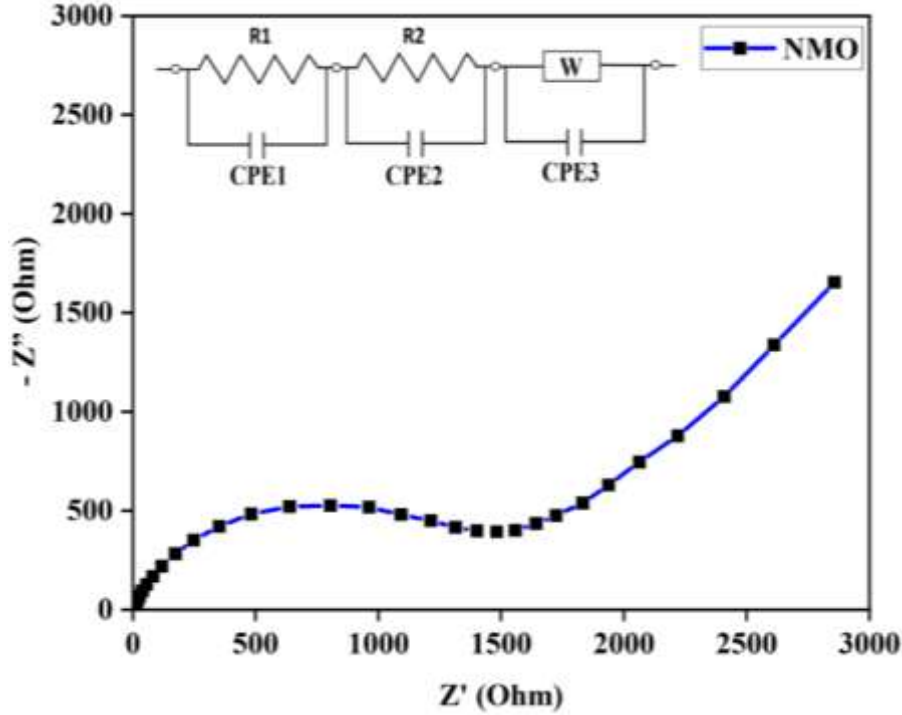


Fig. 17: Nyquist plot showing impedance of NMO electrode conducted at 1 MHz to 100 Hz frequency range and 5 mV AC amplitude.

4.3.2 CV analysis

The CV analysis has been done for further clarification of diffusion coefficient. Fig. 18 shows the CV results of synthesised NMO. The first cathodic scan shows peak attributes to the reduction of NiMn_2O_4 to metallic Nickel and Manganese and formation of Li_2O as shown in equation (7). The subsequent anodic scan attributed to Ni and Mn oxidation as their reaction with Li_2O as shown in reaction (8) and (9), consistent with the overall conversion reaction happens in NMO as an anode material.

The specific capacitance is estimated by using the equation (14):

$$C_p = \frac{A}{2mK(v_2 - v_1)} \quad (14)$$

where, C_p , A , K , and (v_2-v_1) represents specific capacitance, surface area, mass, scan rate and potential window, respectively.

The Randles-Ševčík equation (15) was employed to calculate the diffusion coefficient of Li^+ ions:

$$I_p = 2.69 \times 10^5 * n^{\frac{3}{2}} * A * C_L * \sqrt{D_L} \quad (15)$$

$$D_L = \left(\frac{I_p}{2.69 \times 10^5 * A * n^{\frac{3}{2}} * C_L * K^{1/2}} \right)^2 \quad (16)$$

where, D_L , I_p , K , n , C_L , and A denotes the lithium-ion diffusion coefficient, anodic and cathodic peak current, scan rate, number of electrons involved, concentration of lithium ions, and surface area, respectively. The lithium-ion diffusion coefficient was found to be $10^{-13} - 10^{-14} \text{ cm}^2 \text{ s}^{-1}$. The derived D_L from EIS closely matches with the values obtained through CV analysis, reinforcing the consistency of the diffusion mechanism in NMO using Warburg impedance of the Nyquist plot.

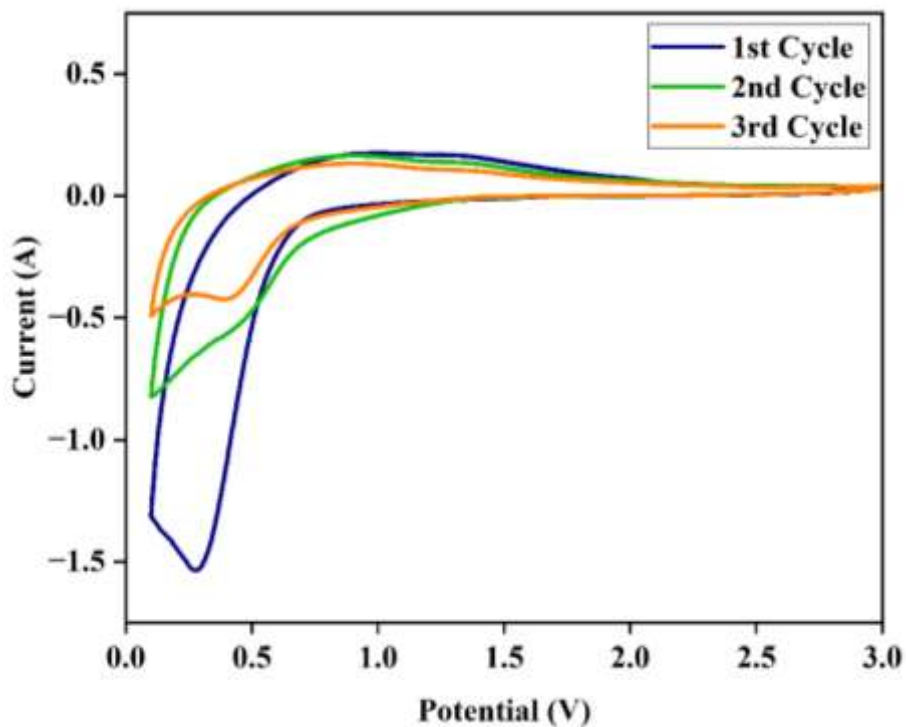


Fig. 18: Cyclic voltammetry curves of synthesised NMO material obtained at 0.1 mV s^{-1} scan rate.

4.3.3 GCD analysis

The specific capacitance along with charge/dis-charge behaviour NMO electrode is evaluated using GCD technique. Figure 19 shows the GCD curves indicating lithium-ion storage behaviour. The redox peaks observed in CV near 0.3-0.5 V, correspond well with the voltage plateaus appeared around 0.6-0.3 V, in the GCD profile, confirming the reversibility of the redox processes. This indicates the same electrochemical process is responsible for both features. The specific capacitance of the electrode has been measured using given formula (17):

$$C_p = \frac{I_m \cdot \Delta t}{\Delta v} \quad (17)$$

where, C_p , I_m , Δt , and Δv represents specific capacitance, current density, discharge time, and potential window, respectively.

The specific capacitance values calculated from GCD were found to be in good agreement with those obtained via CV, supporting the validity of the electrochemical performance.

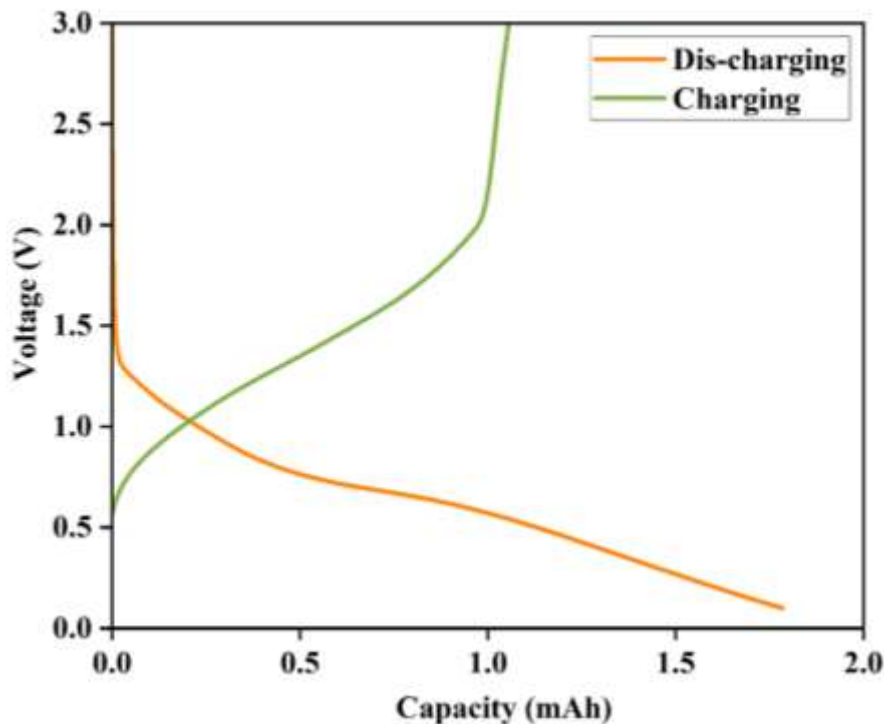


Fig. 19: GCD plot showing charge and discharge capacity of NMO anode material at 0.01C-rate.

Chapter 5

CONCLUSIONS

5.1 Conclusions

A solid-state reaction method was used to synthesize NMO. The XRD results revealed that prepared sample crystallizes in a cubic structure of $Fd3m$ space group, and it matches perfectly with the standard database file no. JCPDS 01-071-0852. The average crystallite size was determined by Scherrer's equation and W-H plot and was found to be ~ 34 nm. The SEM and TEM analysis confirms the hexagonal-shape morphology and polycrystalline nature of synthesised NMO. The particle size evaluated with SEM was found to be 282.32 nm and with TEM was around 180.52 nm. The SAED results correlates with the XRD data for (111), (311), (531), and (731) lattice planes. The ionic conductivity of synthesised electrode was $\sim 1.32 \times 10^{-3} \text{ S cm}^{-1}$. The lithium-ion diffusion coefficient (D_{Li}) for the half-cell, as determined by CV ranged from 10^{-13} to $10^{-14} \text{ cm}^2 \text{ s}^{-1}$, while EIS measurements from Warburg impedance yielded values around $10^{-14} \text{ cm}^2 \text{ s}^{-1}$. Also, the specific capacitance is around 290.02 F g^{-1} , which was analysed and correlated using CV and GCD plots.

5.2 Future Scopes

Further research will focus on synthesizing $\text{NMO}@ \text{MoS}_2$ composites to enhance Li^+ diffusion kinetics. Subsequent Cu-doping of the $\text{NMO}@ \text{MoS}_2$ within this composite will further optimize ion transport and electrode stability. Comprehensive electrochemical analyses will precisely quantify the improved diffusion coefficients in these advanced structures. This work aims to establish a clear correlation between material modification and enhanced Li^+ diffusion for superior battery performance.

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APPENDICES

List of Publications:

- Mehra, S., Jadaun, S.S., Kumar, A., and Panwar, A.K.: **Study of lithium-ion diffusion coefficients through analysis of CV and EIS for NiMn₂O₄ alternate anode material in Lithium-ion batteries.** Revision submitted (under review) at *Journal of Electronic materials*, Springer (IF: 2.2)
- Mehra, S., Yadav, P., and Kumar, A.: **Role of Porphyrins and Corroles for Batteries: A Review.** Under review at *Chemistry Select*, Wiley. (IF: 1.9)
- Mehra, S., Kumar, A., and Panwar, A.K.: **A Review on Bio-Recycling of Lithium: A sustainable Alternative for LIB waste management.** Under communication.
- Mehra, S., Jadaun, S.S., Kumar, A., and Panwar, A.K.: **Investigation of Diffusion Coefficients using Electrochemical Techniques of Anode material composite: NiMn₂O₄@MoS₂.** Under process.

Published:

- Mehra, S. and Gupta, A.: **Green nanotechnology: Path for clean and safer molecular manufacturing.** *Journal of Nanoscience Nanoengineering and Applications*, 12(3), 28-32 (2022) <https://doi.org/10.37591/jonsnea.v12i3.1370>. (IF: 6.004)

List of Conference Attended:

- “International Conference on Advance Materials for Sustainable Future” (ICAMSF-2025), held on 28-29 March 2025 at Chitkara University, Chandigarh, India. (Best Presentation Award)



- 3rd International Conference on “Advanced Functional Materials and Devices” (AFMD-2025) held on 3-5 March 2025 at Atma Ram Sanatan Dharma College, University of Delhi, India.



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